

# SYNTHESIS OF NANO SIZED MESOPOROUS $\gamma$ - $\text{Al}_2\text{O}_3$ POWDER FROM DOMESTIC HAMEDAN KAOLIN

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**Abstract:** . A method for producing high surface area nano-sized mesoporous alumina from inexpensive Iranian kaolin as raw material is proposed. In this method, first, kaolin was purified; for purifying Kaolin, High Grade Magnetic Separation and leaching with HCl and chemical bleaching treatment by using sodium dithionite ( $\text{Na}_2\text{S}_2\text{O}_4$ ) as reducing agent in acidic media ( $\text{H}_2\text{SO}_4$ ) were used. Purified kaolin was calcined. After that, Al (hydr) oxide from acid - leachates of calcined kaolin was precipitated with ammonia, in presence of polyethylene glycol. Finally, a white powder of nano-sized alumina particles was obtained after calcination. BET surface area, X-ray diffraction (XRD), Fourier transform infrared spectroscopy (FT-IR) and scanning electron microscopy (SEM) were used to characterize the sample. The resulting alumina with relatively high surface area ( $201.53 \text{ m}^2 \text{ g}^{-1}$ ) and narrow mean pore diameter (6.91 nm), consists of a particle size distribution ranging from 22 to 36 nm.

**Keyword:** Nano- alumina, kaolin, Synthesis, precipitation.

## 1. INTRODUCTION

Nano  $\text{Al}_2\text{O}_3$  has immense technological and industrial applications, and has been widely used as catalysts [1, 2], catalytic supports [3], abrasive and thermal wear coatings and adsorbents [4-7]. Various methods like sol-gel, hydrothermal process and controlled precipitation of boehmite from solutions of aluminum salts, alkoxides and metal powders have been used for synthesizing nano-alumina. Aluminum alkoxides and other raw materials for processing of alumina are high cost. In recent years, cheaper materials like urban waste [8], low cost Al and  $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$  powders [9] have been used for synthesis of nano alumina. Furthermore, there are inexpensive minerals that are known as a good source of aluminium and can be used for synthesis of nano alumina. Some of these minerals are known as bauxite, kyanite, alunite, sillimanite, andalusite, kaolin, mica and fly ash. Among these minerals, Bauxite is a good source of alumina, but is very scarce. The rareness of bauxite, suggests abundant kaolin as an alternative raw material for production of alumina.

In this paper, kaolin, which is an abundant and

inexpensive mineral in Iran, was used as a starting material for Synthesis of Nano alumina Powder.

Before the leaching, kaolinitic clay must be purified. In a general way, in the state of the art a number of techniques for purifying kaolin clay are known. Among these, the most common comprise: magnetic separation [10], leaching [11, 12], chemical bleaching [13], flotation [14, 15] and selective flocculation [16-18]. The studies show that the calcination of purified kaolin before the leaching is necessary for solubilizing aluminium in HCl acid solution [19]. It has been reported that, Metakaolin is more reactive than kaolin under chemical treatments. After calcination,  $\text{AlO}_6$  octahedrals are transformed into much more reactive tetra and penta coordinated units. The best conditions for obtaining more reactive metakaolins have been discussed by different authors who reported values between 600 and 800 °C [20].

For leaching alumina from kaolin, some of the important advantages of using hydrochloric acid instead of other acids are the insolubility of titanium dioxide and the ease of filtration of slurries which is present in kaolin. Surfactants as

structure-directing agents have been applied to control the pore size, shape and to direct the formation of mesophases on the bases of electrostatic and hydrogen bonding interactions[21]. In this study, the long-chain, inexpensive and non-ionic surfactant-polyethylene glycol (PEG, molar mass 4500) was used as structure-directing agent before the precipitation process. After leaching process, aluminium ions in solution can readily be precipitated as  $\text{Al}(\text{OH})_3$  by adjusting the solution to be neutral or slightly alkaline[22]. The precipitation process was conducted in an aqueous media at room temperature using ammonia. The amorphous alumina powder can be transformed to  $\gamma$ - alumina and  $\delta$  -alumina powder at 700 °C and can be converted to  $\alpha$ -alumina powder above 1100 °C. Researchers demonstrated that, Alumina prepared by the sol-gel method usually exhibit higher pore volumes and surface areas as compared with precipitation in aqueous solutions counterparts[22, 23].

## 2. EXPERIMENTAL

### 2. 1. Materials

All chemicals used are analytical reagents. The kaolin used came from Hamedan province, west of Iran. The iron and calcium content of this kaolin were high.

### 2. 2. Purification of Kaolin

After leaching with water, high energy magnetic separation was done for removing the impurities having magnetic susceptibility such as hematite, titanium, rutile, mica and pyrite. About 47% of iron was removed in this step. For reducing CaO in the sample, the kaolin was leached with hydrochloric acid (7%) at room temperature under 0.5 h stirring. To remove remained iron and other impurities, the kaolin is directed to a chemical bleaching treatment carried out by using sodium dithionite ( $\text{Na}_2\text{S}_2\text{O}_4$ ) as reducing agent in acidic media (pH was adjusted to 3 with sulfuric acid) at 80 °C for 0.5 h[12]. Schematic for the preparation of pure kaolin has been shown in fig. 1.

### 2. 3. Synthesis

First, the purified kaolin was calcined into metakaolin at heating rate of 10 °C/min and the sample was kept for 2.5 h at the required temperatures to loosen the alumina component. Chemical composition of Kaolin, purified Kaolin and metakaolin (calcined at 700°C) are shown in table 1. Then, the metakaolin was leached with concentrated hydrochloric acid (6 M) at 90 °C under 3 h stirring, then; suspensions were filtered and the filtrate was collected for producing of aluminum hydroxide. After that, following addition of polyethylene glycol (PEG, molar mass 4500), ammonia (.75 M) was added. The precipitated Al (hydr) oxide was filtrated, washed with de-ionized water and dried. Finally, Subsequent calcination yielded nano-sized  $\text{Al}_2\text{O}_3$ . Schematic for the preparation of alumina nanoparticles from purified kaolin has been shown in Fig. 2.

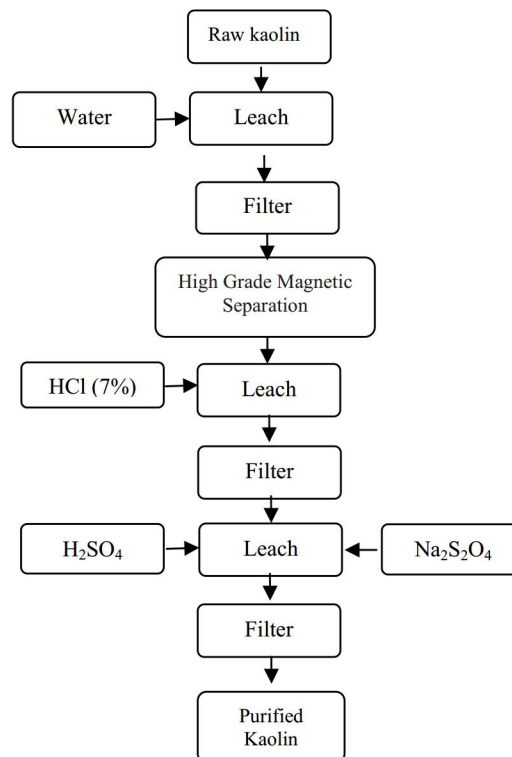


Fig.1. Schematic for the preparation of pure kaolin

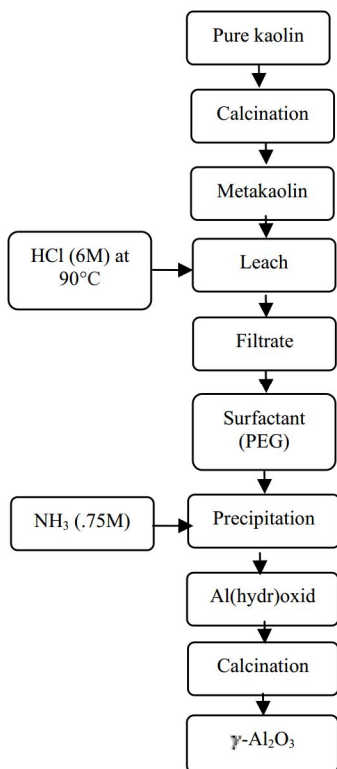


Fig.2. Schematic for the preparation of nano  $\gamma$ - alumina from kaolin

## 2. 4. Characterization

A BET surface area was used to determine the surface area of the sample, and scanning electron microscopy (SEM) was used to determine the particle size and morphology. Sample phases were identified using X-ray diffractometry (XRD) with nickel-filtered Cu K radiation.

## 3. RESULTS AND DISCUSSION

### 3.1. Activation and Subsequent Leaching of Raw Kaolin

The XRD patterns of purified kaolin and calcined kaolin in 600 and 800 °C are shown in Fig. 3. With increasing temperature of calcinations, gradual disappearance of the kaolinite reflections has been seen and kaolin almost transformed into metakaolin. The studies showed that the calcination of kaolin prior to the leaching is necessary to solubilize aluminium in

Table 1. Chemical composition of Kaolin and Metakaolin (calcined at 700°C)

Substance	Raw Kaolin	Purified Kaolin	Metakaolin
Na <sub>2</sub> O	1.292	0.117	0.131
MgO	1.212	0.113	0.119
Al <sub>2</sub> O <sub>3</sub>	29.77	35.366	44.958
SiO <sub>2</sub>	42.89	46.216	50.598
P <sub>2</sub> O <sub>5</sub>	0.297	0.166	0.148
Cl	0.189	-	-
SO <sub>3</sub>	0.389	0.078	0.051
TiO <sub>2</sub>	1.83	0.471	0.527
Fe <sub>2</sub> O <sub>3</sub>	1.79	0.366	0.371
CaO	0.841	0.058	0.055
LOI	19.5	17.01	3.04

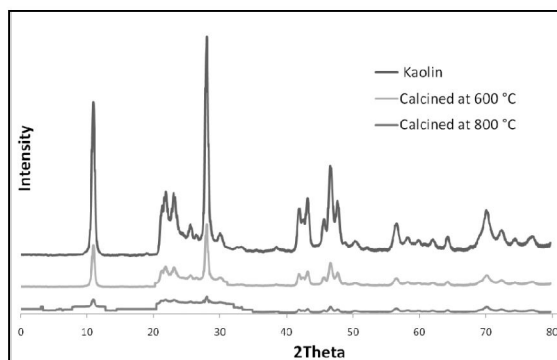


Fig. 3. XRD patterns of purified and calcined kaolin at 600 and 800°C.

HCl acid solution. It was reported that the calcination should be realized at 750 °C for at least 2 h[19]. It was also claimed that after the calcination process and converting kaolinite (Al<sub>2</sub>Si<sub>2</sub>O<sub>5</sub> (OH)<sub>4</sub>) to metakaoline (Al<sub>2</sub>Si<sub>2</sub>O<sub>7</sub>), aluminium can be extracted much more rapidly and easily [24]. Pure kaolin contains 35.36% Al<sub>2</sub>O<sub>3</sub>. After calcination, this ratio increases to 44.96% since all water in the structure is removed. A known amount of sample was digested with 6 M HCl acid solution at 90 °C for 3 h under a reflux condenser. The leaching rate of Al<sup>3+</sup> ions from raw kaolin and metakaolin is shown in Tab. 2. Only a few amounts of Al<sup>3+</sup> ions (about 6.4%) were leached from purified kaolin. With acid treatment, Fe<sup>3+</sup>, Ti<sup>4+</sup> (lower than others) as rutile and Si<sup>4+</sup> as SiO<sub>2</sub> were leached and these materials can improve the thermal stability of the alumina to some degree.

**Table 2.** Leaching rate of Al<sup>3+</sup> ions from purified kaolin and metakaolins

Temperature (°C)	purified kaolin	500	600	800
Leaching rate(%)	6.4	58.3	82.7	89.2

The sample calcined at 900 °C showed an Al<sup>3+</sup> ions leaching rate of about 82.9%. The lower reactivity after calcination at still higher temperatures may be related to the sintering of kaolin particles at 900 °C and to the gradual transformation of metakaolin into mullite[20].

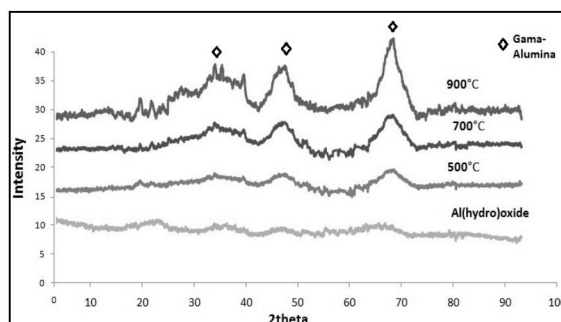
### 3. 2. Formation of Nano Alumina

By removal of the adsorbed water, small mass loss beneath 114 °C was seen. Due to the removal of water during the formation of the oxide phases, The mass loss of the Al Between 500 and 900 °C, the three main reflections of nano  $\gamma$  - Al<sub>2</sub>O<sub>3</sub> phase are obviously observed as broad peaks at 2 angles around 38.0°, 46.0°, and 66.0° which correspond to the [3 1 1], [4 0 0], and [4 4 0] planes, respectively. The peaks in all patterns are meaningfully broad indicating the formation of nano sized  $\gamma$  - Al<sub>2</sub>O<sub>3</sub> crystallites. XRD patterns of the Al (hydr) oxide and calcined samples at different temperatures are shown in Fig. 4. The crystallite sizes were calculated by using Scherrer equation:

$$(1)$$

Where  $\lambda$  is the wavelength of the X-rays,  $\Delta 2\theta$  is the full width of diffraction peak at half maximum intensity and  $\theta$  is the Bragg angle. The calculated crystallite sizes were found to be in the range of 22–36 nm.

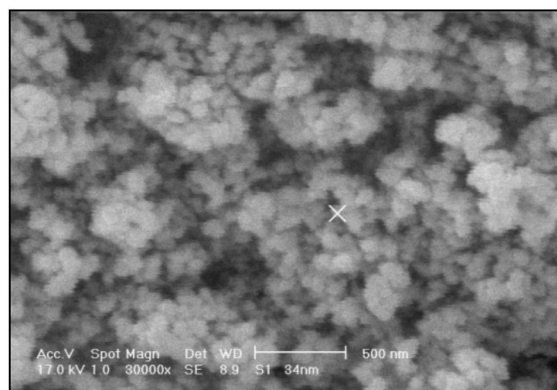
As can be seen from Fig. 4, the precursor Al (hydr) oxide show amorphous structure and with increasing temperature of calcination, the crystal phase increased and the amorphous Al (hydr) oxide have almost transformed into  $\gamma$ -Al<sub>2</sub>O<sub>3</sub>. The addition of the surfactants resulted in the



**Fig. 4.** XRD patterns of the Al (hydr) oxide and calcined samples at different temperatures.

formation of an amorphous Al (hydr) oxide phase coupled with some hydrates of surfactant. This may be the result of the adsorption of surfactant layer on the surface of the initial Al (hydr) oxide nuclei, which prevented the aggregation and grain growth of Al (hydr) oxide in the absence of surfactants that would allow the formation of crystalline phase. Fig. 5 shows the SEM micrograph for  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> powders. The  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> indicated low agglomeration of particles with uniform sizes and spherical in shape.

According to FTIR spectra (Fig. 6) the bands in the region of 400–1000 cm<sup>-1</sup> are generally associated with the stretching vibration of Al–O bonds[25] and broad bands around 3500 cm<sup>-1</sup> and 1630 cm<sup>-1</sup> are assigned to stretching and bending modes of adsorbed water. The bands at 500–750 cm<sup>-1</sup> are assigned to  $\nu$ -AlO<sub>6</sub>, whereas



**Fig.5.** SEM graph of nano  $\gamma$  -alumina

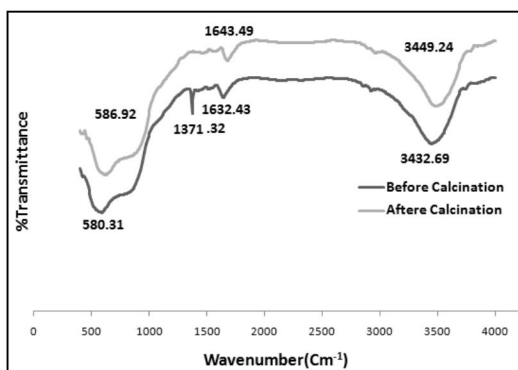


Fig. 6. FT- IR spectra of the nano  $\gamma$ -alumina before and after calcination at 700 °C.

the band around 900  $\text{cm}^{-1}$  corresponded to  $\nu$ - $\text{AlO}_4$ [23]. The bands from 1400  $\text{cm}^{-1}$  to 1600  $\text{cm}^{-1}$  indicated the formation of alumina[26]. The composition of synthesised Nano  $\gamma$  -  $\text{Al}_2\text{O}_3$  is shown in Tab. 3. The absence of band around 1371  $\text{cm}^{-1}$  due to C–H implied that the alumina was without PEG after calcination.

$\text{N}_2$  adsorption/desorption experiments for synthesised Nano  $\gamma$ -  $\text{Al}_2\text{O}_3$  were carried out using a Belsorp mini II (BelJapan), and pore size distributions were calculated using the Barret-Joyner-Halenda (BJH) model on the desorption branch.

Synthesised Nano  $\gamma$ -  $\text{Al}_2\text{O}_3$  showed a narrow pore size distribution with a high surface area (201.53  $\text{m}^2 \text{g}^{-1}$ ), a mean pore diameter of 6.91 nm

Table 3. Composition of synthesised Nano  $\gamma$  -  $\text{Al}_2\text{O}_3$

Substance	Nano $\gamma$ - $\text{Al}_2\text{O}_3$
MgO	0.11
$\text{Al}_2\text{O}_3$	90.021
$\text{SiO}_2$	0.376
$\text{P}_2\text{O}_5$	0.188
Cl	0.233
$\text{SO}_3$	0.027
Zn	0.24
$\text{Fe}_2\text{O}_3$	0.648
CaO	0.087
Pb	0.029
Ga	0.011
LOI	8.03

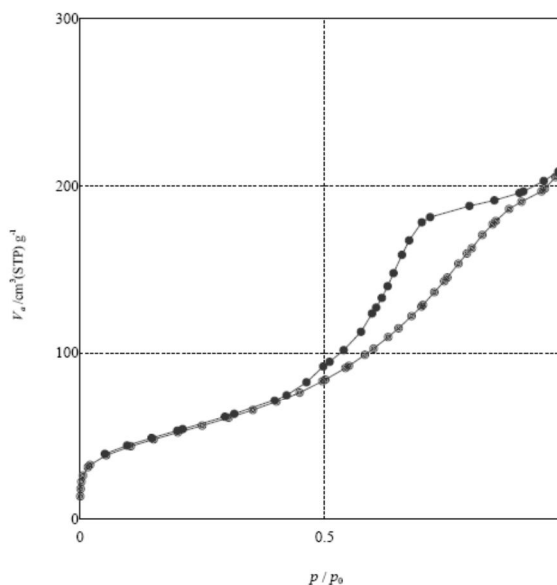


Fig. 7.  $\text{N}_2$  adsorption/desorption isotherm for synthesised Nano  $\gamma$  -  $\text{Al}_2\text{O}_3$ .

and a high pore volume (0.33  $\text{cm}^3 \text{g}^{-1}$ ). The location of the hysteresis loop in the  $\text{N}_2$  isotherm obtained nano  $\gamma$ -  $\text{Al}_2\text{O}_3$  sample (Fig. 7) displays type IV hysteresis pointing out the presence of mesoporosity and can be used to determine whether the material possessed a regular framework pore or interparticle voids, such as a textural pore. The framework porosity at 0.4-0.75  $P/P_0$  in the  $\text{N}_2$  isotherm shows that the porosity was framework, while the textural porosity at 0.8-1  $P/P_0$  indicates porosity arising from the noncrystalline intra-aggregate voids and spaces formed by interparticle contacts. The large surface areas (201.53 ) and narrow pore size distributions (around 7 nm) integrated with excellent thermal stability promised potential applications of these mesoporous nano  $\text{Al}_2\text{O}_3$  in catalysis [1, 27] and water treatment applications[5, 28, 29].

### 3. 3. Formation Mechanism

After leaching of metakaolin, in presence of PEG (mass weight of 4500) as a surfactant and as dispersant for preventing the solid particles from aggregation, Al (hydr) oxide was precipitated. The reason for using PEG is that PEG can

obstruct the hydrolysis and condensation of the formed Al (hydr) oxide particles, e. g. be adsorbed on the surface of the growing particles[30]. In the near neutral conditions, PEG molecules adsorbed on the surfaces of aluminum hydrate would limit its growth during the precipitation process and then lead to the formation of colloid particles.

#### 4. CONCLUSION

A simple method for producing Nano sized mesoporous  $\gamma$  -  $\text{Al}_2\text{O}_3$  from impure Hamedan kaolin with spherical shape, particle size distribution ranging from 22 to 36 nm with relatively high surface area ( $201.53 \text{ m}^2 \text{ g}^{-1}$ ) and narrow mean pore diameter (6.91 nm), is described. In this method pre calcination of kaolin plays an important role in leaching of aluminium in solution. Besides, addition of surfactant before precipitation of nano alumina was used as dispersant for preventing the nano particles from aggregation.

#### 5. ACKNOWLEDGMENTS

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